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**Modern Chemistry Chapter 14
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Modern Chemistry Chapter 14 Acids & Bases Section 1 Properties of Acids & Bases Aqueous solutions of acids taste sour. Acids change the color of acid-base indicators. Some acids react with active metals and release hydrogen gas. Acids react with bases to produce salts and

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water. Acids conduct electricity (electrolytes).

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84 Name Section Quiz, continued Class

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CHAPTER 14 REVIEW . Acids and Bases.
SHORT ANSWER Answer the following
questions in the space provided. 1. a.
Write the two equations that show the
two-stage ionization of sulfurous acid in

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water. stage 1: $\text{H}_2\text{SO}_3(\text{aq}) + \text{H}_2\text{O}(\text{l}) \sim$
 $\text{H}_3\text{O}^+(\text{aq}) + \text{HSO}_3^-(\text{aq})$ b. Which stage of
ionization usually produces more ions?
Explain your answer.

CHAPTER 14 REVIEW Acids and Bases

CHAPTER 14 REVIEW Acids and Bases
SECTION 2 SHORT ANSWER Answer the

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following questions in the space provided. 1. a. Write the two equations that show the two-stage ionization of sulfurous acid in water. stage 1: $\text{H}_2\text{SO}_3(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{H}_3\text{O}^+(\text{aq}) + \text{HSO}_3^-(\text{aq})$
stage 2: $\text{HSO}_3^-(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{H}_3\text{O}^+(\text{aq}) + \text{SO}_3^{2-}(\text{aq})$ b.

14 Acids and Bases

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CHAPTER 14 REVIEW Ions in Aqueous Solutions and Colligative Properties
SECTION 14-1 SHORT ANSWER Answer the following questions in the space provided. 1. Use the guidelines in Table 14-1 on page 427 of the text to predict the solubility of the following compounds in water: a. magnesium nitrate b. barium sulfate c. calcium carbonate d.

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ammonium phosphate

14 Ions in Aqueous Solutions and Colligative Properties

Zumdahl's Chemistry, 5th Edition
Textbook; Holt Earth Science Chapter
12, Section 12.3; Holt Earth Science
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Modern Chemistry Chapter 15 Acid-Base
Titration & pH Section 1 self-ionization of
water occurs when two water molecules
produce a hydronium (H_3O^+) and a

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hydroxide (OH⁻) ion 2 H₂O H₃O⁺ + OH⁻
The ionization constant (K_w) of water is:
 $K_w = [H_3O^+] [OH^-] = 1.0 \times 10^{-14} \text{ M}$
Acidic, Basic, & Neutral IF $[H_3O^+] >$
 $[OH^-]$ then solution is acidic.

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16.1-16.3. White Board Practice: CH 14
Naming Acids.

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Chemistry Is a Physical Science 3 ...
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LA.PS.14 Identify unknowns as elements, compounds, or mixtures based on physical properties (e.g., ...

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Section 1. Gases & Pressure. Pressure &
Force. ... $1.0 \text{ atm} = 14.700 \text{ psi}$. Do
practice problems #1 & #2 on page 365

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Modern Chemistry 6 The Periodic Law
CHAPTER 5 REVIEW The Periodic Law
SECTION 3 SHORT ANSWER Answer the
following questions in the space
provided. 1. _____ When an electron is

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added to a neutral atom, energy is (a)
always absorbed.

Modern Chemistry Chapter 5 Review Answers Section 1

Modern Chemistry 96 Quiz Section Quiz:
Aqueous Solutions and the Concept of
pH In the space provided, write the
letter of the term or phrase that best

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completes ... 14 Acids and Bases

Section: Properties of Acids and Bases 1.

d 2. c 3. b 4. a 5. a 6. 7. a 8. c 9. d 10. b

Section: Acid-Base Theories 1. c 2. b 3. a

4. b 5. b 6. a 7. d 8. c

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