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Pearson Edexcel GCE in Chemistry (6CH05) Paper 01 General Principles of Chemistry II – Transition ... Section A (multiple choice) Question Number Correct Answer Mark 1 1. The only correct answer is C A is not correct because neither Al nor H has oxidation number +5 B is not correct because neither K nor Mn has

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EDEXCEL INTERNATIONAL GCSE CHEMISTRY EDEXCEL CERTIFICATE IN CHEMISTRY ANSWERS SECTION D Chapter 18 1. a) (i) methane (ii) propane (iii) pentane (iv) propene (v) ethene (vi) but-1-ene b) 2. a) The existence of molecules with the same molecular formula but different structural formulae. b)

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Chemistry EDEXCEL INTERNATIONAL GCSE CHEMISTRY EDEXCEL CERTIFICATE IN CHEMISTRY ANSWERS SECTION A Chapter 1 1. a) melting b) freezing c) subliming / sublimation d) subliming / sublimation 2. a) Note: Solids should have regularly packed particles touching. Liquids should have most of the

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Pearson Edexcel GCE in Chemistry (6CH04) Paper 01 General Principles of Chemistry I - Rate, Equilibria and Further ... Section A (multiple choice) Question Number Answer Mark 1 1. The only correct answer is C A is not correct because colour increases as bromine is formed B is not correct because colour increases as bromine is formed

Mark Scheme (Results)

EDEXCEL INTERNATIONAL GCSE CHEMISTRY EDEXCEL CERTIFICATE IN CHEMISTRY ANSWERS SECTION C Chapter 16 1. Burn sulfur in air to give sulfur dioxide. $S(s) + O_2(g) \rightarrow SO_2(g)$ Pass this with more air over a vanadium(V) oxide catalyst at a temperature of 450°C and a pressure of 1 to 2 atmospheres to give sulfur trioxide.

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explored in some detail in chemistry at a higher level. b) Correct electronic structures for Al 2,8,3 and F 2,7. Diagrams (similar to those in the chapter) showing transfer of electrons, and the charges and electronic structures of the ions formed. (or words to the same effect).

IGCSE Chemistry A answers

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GRADEGORILLA | Chemistry Revision Questions

Answers Edexcel International GCSE Chemistry Teacher Pack 5 © HarperCollinsPublishers Ltd 2012
Page 46 1. 1 mole of magnesium (24 g) produces 1 mole of hydrogen gas (24 000 cm³). So 1/6 mole of Mg produces 1/6 mole hydrogen = 4000 cm³ at room temperature and pressure. 2. 1 mole of oxygen gas is produced from 2 moles of hydrogen peroxide.

Answers Section 1. Principles of chemistry

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